

Name _____

Chem 1194: K_{sp} Determination Report

155

Date _____ Section _____

Your Data

Exact concentration of ~0.1 M HCl used _____

Exact concentration of ~0.04 M HCl used _____

Mass of $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ used _____

Solution Titrated	Initial appearance of the solutions titrated. (clear or turbid)	Volume HCl / mL Mean \pm SD	$[\text{OH}^-]$ / M	$\text{Ca}(\text{OH})_2$ solubility (in g/100 mL)
Sat'd $\text{Ca}(\text{OH})_2$ filtrate/decantate (circle one)		\pm		
$\text{Ca}(\text{OH})_2$ + $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ filtrate/decantate (circle one)		\pm		

	$[\text{Ca}^{2+}]$ (M)	K_{sp}	S_K	\mathcal{K}_{sp}
Sat'd $\text{Ca}(\text{OH})_2$				
$\text{Ca}(\text{OH})_2$ + $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$				

Constancy of K_{sp} versus Solubility (from your two solutions)Relative difference in $\text{Ca}(\text{OH})_2$ solubility _____ %Relative difference of K_{sp} _____ %Relative difference of \mathcal{K}_{sp} _____ %