

Name _____

Chem 1194: Kinetics Report

99

Date _____ Section _____

Group Number ____

Others in Your Group

Last Name and (First Name or Initial)

Group Runs

T_{final} of Run ($^{\circ}\text{C}$) _____

$1/T_{\text{final}}$ of Run (K^{-1}) _____

Time of Run (s) _____

$\ln(t)$ of Run _____

Using the data in the table above, submit an Excel graph of $\ln t$ as a function of $1/T$. (Your own - *not* a copy of your group's.) On the Excel graph, report the trendline with units.

Sample Calculation of E_a from the slope and R.

Prediction of Reaction Time at "Common" Temperature

Given temperature _____ Predicted $\ln t$ _____, time _____

Sample calculation of these

Calculation of relative difference(%)

Calculation of Your
Group's $[\text{I}_2]_{\text{initial}}$

Calculation of Your
Group's Reaction Rate

Determination of Rates for Each Group at "Common" Temperature

Group #	[I ₂] _{initial} (M)	[acetone] _{initial} (M)	[H ⁺] _{initial} (M)	Time (s)	Rate (mole rxn/L s)
1	_____	_____	_____	_____	_____
2	_____	_____	_____	_____	_____
3	_____	_____	_____	_____	_____
4	_____	_____	_____	_____	_____

_____	_____	_____
$\ln([I_2]_{\text{group1}}/[I_2]_{\text{group2}})$	$\ln([\text{acet}]_{\text{grp1}}/[\text{acet}]_{\text{grp3}})$	$\ln([HCl]_{\text{grp1}}/[HCl]_{\text{grp4}})$

Reaction Orders from Relative Rates at Common Temperature

$$\frac{\text{Rate}_1}{\text{Rate}_2} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_2} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_2)}{\ln([I_2]_1/[I_2]_2)} = \text{_____} \quad m = \text{_____}$$

$$\frac{\text{Rate}_1}{\text{Rate}_3} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_3} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_3)}{\ln([\text{acet}]_1/[\text{acet}]_3)} = \text{_____} \quad n = \text{_____}$$

$$\frac{\text{Rate}_1}{\text{Rate}_4} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_4} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_4)}{\ln([HCl]_1/[HCl]_4)} = \text{_____} \quad p = \text{_____}$$

Write the rate law: Rate = _____

Common Temperature Rate Constant

	Group 1	Group 2	Group 3	Group 4
Rate Constant, k (_____) unit	_____	_____	_____	_____

_____ Sample calculation
Average rate constant, k of k for your group