

Date \_\_\_\_\_ Section \_\_\_\_\_

Group Number \_\_\_\_

Others in Your Group

Last Name and (First Name or Initial)

**Group Runs** $T_{\text{average}}$  of Run ( $^{\circ}\text{C}$ )    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_ $1/T_{\text{average}}$  of Run ( $\text{K}^{-1}$ )    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_

Time of Run (s)    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_

 $\ln(t)$  of Run    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_    \_\_\_\_\_

Sample Calculation

of  $T_{\text{average}}$  from $\frac{\text{Iodine T}}{\text{Acetone T}} = \frac{\text{Final T}}{\text{Final T}}$ 

Using the data in the table above, submit an Excel graph of  $\ln t$  as a function of  $1/T$ . (Your own - *not* a copy of your group's.) On the Excel graph, report the trendline with units.

Sample Calculation of  $E_a$  from the slope and R.**Prediction of Reaction Time at "Common" Temperature**Given temperature \_\_\_\_\_ Predicted  $\ln t$  \_\_\_\_\_, time \_\_\_\_\_

Sample calculation of these

Calculation of relative difference(%)

Calculation of Your  
Group's  $[I_2]_{\text{initial}}$ Calculation of Your  
Group's Reaction Rate

**Determination of Rates for Each Group at “Common” Temperature**

Group #	$[I_2]_{\text{initial}} / M$	$[\text{acetone}]_{\text{initial}} / M$	$[H^+]_{\text{initial}} / M$	Time /s	Rate /(mole rxn/L s)
1	_____	_____	_____	_____	_____
2	_____	_____	_____	_____	_____
3	_____	_____	_____	_____	_____
4	_____	_____	_____	_____	_____

$$\ln([I_2]_{\text{group1}}/[I_2]_{\text{group2}}) \quad \ln([\text{acet}]_{\text{grp1}}/[\text{acet}]_{\text{grp3}}) \quad \ln([HCl]_{\text{grp1}}/[HCl]_{\text{grp4}})$$

**Reaction Orders from Relative Rates at Common Temperature**

$$\frac{\text{Rate}_1}{\text{Rate}_2} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_2} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_2)}{\ln([I_2]_1/[I_2]_2)} = \text{_____} \quad m = \text{_____}$$

$$\frac{\text{Rate}_1}{\text{Rate}_3} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_3} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_3)}{\ln([\text{acet}]_1/[\text{acet}]_3)} = \text{_____} \quad n = \text{_____}$$

$$\frac{\text{Rate}_1}{\text{Rate}_4} = \text{_____} \quad \ln \frac{\text{Rate}_1}{\text{Rate}_4} = \text{_____} \quad \frac{\ln(\text{Rate}_1/\text{Rate}_4)}{\ln([HCl]_1/[HCl]_4)} = \text{_____} \quad p = \text{_____}$$

**Write the rate law:** Rate = \_\_\_\_\_**Common Temperature Rate Constant**

	Group 1	Group 2	Group 3	Group 4
Rate Constant, $k$ ( _____ unit )	_____	_____	_____	_____

\_\_\_\_\_ Sample calculation  
Average rate constant,  $k$  of  $k$  for your group